Chapter 7 Review

Pages 266 - 267 (1-10, 14, 21, 22)
Pages 268-269 (1-14 advanced) (1-10 regular)

Pages 266-267

1. B
2. C
3. A
4. A
5. C
6. B
7. C
8. D
9. C
10. D

14. Sample Answer: As energy is removed from the gas, its temperature decreases. At its condensation point, the temperature remains constant, and the gas changes to a liquid. The temperature then continues to decrease. At its freezing point, the temperature again remains constant, and the liquid changes to a solid.

21. The density of the ice is 0.92 g/cm^3. It floats in liquid water because its density is less than that of liquid water.
22. Gold has a greater density (19.3 g/cm^3 compared to 11.4 g/cm^3 for lead)

Pages 268-269

1. C DOK 1
2. C DOK 2
3. D DOK 2
4. D DOK 1
5. D DOK 2
6. C DOK 1
7. C DOK 2
8. A DOK 4
9. D DOK 4
10. D DOK 1
11. Part B on the graph shows that the temperature does not increase as a solid melts. As the ice in water melts, the temperature of the water and the ice will stay at the melting point of water, 0 degrees C. DOK 3
12. Parts C and D of the graph show what happens to cold water that is put on a stove to boil. First the temperature of the water will increase as energy is added to it. Then, when the water begins to boil, the temperature stays at the boiling point of water. DOK 4
13. The first step would be to filter the mixture to separate the sand from the water and sugar. The next step would be to boil the mixture to make the water evaporate. The sugar will be left behind. DOK 4
14. The equation is incorrect because it is not balanced. A correct equation would show conservation of mass by showing equal numbers of atoms for each element on both sides of the arrow. The products side shows more hydrogen and chlorine atoms. DOK 3